The Double Bulbed Tube

- Put your hand on one of the bulbs. What does this do to the temperature of that bulb?

- As the temperature of the bulb increases, the volume of gas in the bulb also increases. This pushes the liquid into the other bulb. The gas pressure in each bulb remains the same though.

- Why is there no effect when you hold the connecting tube?

The Ideal Gas Law:

\[
\text{Volume} = \frac{\text{Temperature} \times k}{\text{Pressure}}
\]

\[k = n \times R, \text{ where } n = \text{number of moles of gas}\]

\[R = \text{universal gas constant}\]