STANDARD REDUCTION POTENTIALS IN AQUEOUS SOLUTION AT 25°C

Increasing strength of oxidizing agents ↑

Half-reaction			$E°$(V)
$F_2(g) + 2\,e^-$	→	$2\,F^-$	2.87
$Co^{3+} + e^-$	→	Co^{2+}	1.82
$Au^{3+} + 3\,e^-$	→	$Au(s)$	1.50
$Cl_2(g) + 2\,e^-$	→	$2\,Cl^-$	1.36
$O_2(g) + 4\,H^+ + 4\,e^-$	→	$2\,H_2O(l)$	1.23
$Br_2(l) + 2\,e^-$	→	$2\,Br^-$	1.07
$2\,Hg^{2+} + 2\,e^-$	→	Hg_2^{2+}	0.92
$Hg^{2+} + 2\,e^-$	→	$Hg(l)$	0.85
$Ag^+ + e^-$	→	$Ag(s)$	0.80
$Hg_2^{2+} + 2\,e^-$	→	$2\,Hg(l)$	0.79
$Fe^{3+} + e^-$	→	Fe^{2+}	0.77
$I_2(s) + 2\,e^-$	→	$2\,I^-$	0.53
$Cu^+ + e^-$	→	$Cu(s)$	0.52
$Cu^{2+} + 2\,e^-$	→	$Cu(s)$	0.34
$Cu^{2+} + e^-$	→	Cu^+	0.15
$Sn^{4+} + 2\,e^-$	→	Sn^{2+}	0.15
$S(s) + 2\,H^+ + 2\,e^-$	→	$H_2S(g)$	0.14
$2\,H^+ + 2\,e^-$	→	$H_2(g)$	0.00
$Pb^{2+} + 2\,e^-$	→	$Pb(s)$	−0.13
$Sn^{2+} + 2\,e^-$	→	$Sn(s)$	−0.14
$Ni^{2+} + 2\,e^-$	→	$Ni(s)$	−0.25
$Co^{2+} + 2\,e^-$	→	$Co(s)$	−0.28
$Cd^{2+} + 2\,e^-$	→	$Cd(s)$	−0.40
$Cr^{3+} + e^-$	→	Cr^{2+}	−0.41
$Fe^{2+} + 2\,e^-$	→	$Fe(s)$	−0.44
$Cr^{3+} + 3\,e^-$	→	$Cr(s)$	−0.74
$Zn^{2+} + 2\,e^-$	→	$Zn(s)$	−0.76
$2\,H_2O(l) + 2\,e^-$	→	$H_2(g) + 2\,OH^-$	−0.83
$Mn^{2+} + 2\,e^-$	→	$Mn(s)$	−1.18
$Al^{3+} + 3\,e^-$	→	$Al(s)$	−1.66
$Be^{2+} + 2\,e^-$	→	$Be(s)$	−1.70
$Mg^{2+} + 2\,e^-$	→	$Mg(s)$	−2.37
$Na^+ + e^-$	→	$Na(s)$	−2.71
$Ca^{2+} + 2\,e^-$	→	$Ca(s)$	−2.87
$Sr^{2+} + 2\,e^-$	→	$Sr(s)$	−2.89
$Ba^{2+} + 2\,e^-$	→	$Ba(s)$	−2.90
$Rb^+ + e^-$	→	$Rb(s)$	−2.92
$K^+ + e^-$	→	$K(s)$	−2.92
$Cs^+ + e^-$	→	$Cs(s)$	−2.92
$Li^+ + e^-$	→	$Li(s)$	−3.05

Increasing strength of reducing agents ↓

This table is from http://apcentral.collegeboard.com/apc/members/repository/ap04_chem_standard_9793.pdf.
It has been modified by the addition of the notations showing directions of increasing tendencies towards oxidation and reduction.